



BALANCE THE GIVEN CHEMICAL EQUATIONS

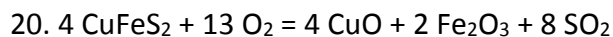
Worksheet - 14

1. $\text{NaHCO}_3 = \text{Na}_2\text{CO}_3 + \text{CO}_2 + \text{H}_2\text{O}$
2. $16 \text{Cr} + \text{S}_8 = \text{Cr}_2\text{S}_3$
3. $4 \text{FeO} + \text{O}_2 = \text{Fe}_2\text{O}_3$
4. $\text{Mg}(\text{OH})_2 + \text{HCl} = \text{MgCl}_2 + 2 \text{H}_2\text{O}$
5. $2 \text{K} + 2 \text{H}_2\text{O} = \text{KOH} + \text{H}_2$
6. $\text{C}_2\text{H}_6 + 7 \text{O}_2 = 4 \text{CO}_2 + \text{H}_2\text{O}$
7. $\text{C}_{10}\text{H}_{22} + 31 \text{O}_2 = \text{CO}_2 + 22 \text{H}_2\text{O}$
8. $10 \text{C}_2\text{H}_6 + \text{O}_2 = 3 \text{H}_2\text{O} + 20 \text{CO}_2$
9. $2 \text{Pb}(\text{NO}_3)_2 = \text{PbO} + 4 \text{NO}_2 + \text{O}_2$
10. $\text{NaOH} + \text{H}_2\text{SO}_4 = \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$
11. $\text{Na}_2\text{S}_2\text{O}_3 + \text{H}_2\text{SO}_4 = \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + 4 \text{SO}_2 + 2 \text{S}$
12. $10 \text{C} + \text{HNO}_3 = \text{CO}_2 + 20 \text{NO}_2 + \text{H}_2\text{O}$
13. $5 \text{Na}_2\text{S}_2\text{O}_3 + \text{KMnO}_4 + 7 \text{H}_2\text{SO}_4 = \text{Na}_2\text{SO}_4 + 8 \text{MnSO}_4 + 4 \text{K}_2\text{SO}_4 + 7 \text{H}_2\text{O}$
14. $\text{Sm}_3\text{Zr}_4\text{P}_6\text{O}_{24} = \text{SmZr}_2\text{P}_3\text{O}_{12} + 8 \text{SmPO}_4 + 8 \text{ZrO}_2 + \text{P}_4$
15. $\text{H}_2\text{SO}_4 + \text{C}_2\text{H}_6 = 2 \text{CO}_2 + \text{SO}_2 + 10 \text{H}_2\text{O}$
16. $\text{FeS} + \text{HCl} = \text{FeCl}_2 + \text{H}_2\text{S}$
17. $\text{Na}_2\text{S}_2\text{O}_3 + 4 \text{HCl} = 4 \text{NaCl} + \text{SO}_2 + 3 \text{S} + 2 \text{H}_2\text{O}$
18. $\text{HClO}_4 + \text{P}_4\text{O}_{10} = \text{H}_3\text{PO}_4 + 6 \text{Cl}_2\text{O}_7$
19. $\text{S}^{2-} + 2 \text{H}^{+} + \text{H}_2\text{O}_2 = \text{S} + \text{H}_2\text{O}$
20. $\text{CuFeS}_2 + 13 \text{O}_2 = 4 \text{CuO} + \text{Fe}_2\text{O}_3 + 8 \text{SO}_2$



ANSWERS

1. $2 \text{NaHCO}_3 = \text{Na}_2\text{CO}_3 + \text{CO}_2 + \text{H}_2\text{O}$
2. $16 \text{Cr} + 3 \text{S}_8 = 8 \text{Cr}_2\text{S}_3$
3. $4 \text{FeO} + \text{O}_2 = 2 \text{Fe}_2\text{O}_3$
4. $\text{Mg}(\text{OH})_2 + 2 \text{HCl} = \text{MgCl}_2 + 2 \text{H}_2\text{O}$
5. $2 \text{K} + 2 \text{H}_2\text{O} = 2 \text{KOH} + \text{H}_2$
6. $2 \text{C}_2\text{H}_6 + 7 \text{O}_2 = 4 \text{CO}_2 + 6 \text{H}_2\text{O}$
7. $2 \text{C}_{10}\text{H}_{22} + 31 \text{O}_2 = 20 \text{CO}_2 + 22 \text{H}_2\text{O}$
8. $10 \text{C}_2\text{H}_6 + 20 \text{O}_2 = 3 \text{H}_2\text{O} + 20 \text{CO}_2$
9. $2 \text{Pb}(\text{NO}_3)_2 = 2 \text{PbO} + 4 \text{NO}_2 + \text{O}_2$
10. $2 \text{NaOH} + \text{H}_2\text{SO}_4 = \text{Na}_2\text{SO}_4 + 2 \text{H}_2\text{O}$
11. $\text{Na}_2\text{S}_2\text{O}_3 + 3 \text{H}_2\text{SO}_4 = \text{Na}_2\text{SO}_4 + 3 \text{H}_2\text{O} + 4 \text{SO}_2 + 2 \text{S}$
12. $10 \text{C} + 20 \text{HNO}_3 = 10 \text{CO}_2 + 20 \text{NO}_2 + \text{H}_2\text{O}$
13. $5 \text{Na}_2\text{S}_2\text{O}_3 + 8 \text{KMnO}_4 + 7 \text{H}_2\text{SO}_4 = 5 \text{Na}_2\text{SO}_4 + 8 \text{MnSO}_4 + 4 \text{K}_2\text{SO}_4 + 7 \text{H}_2\text{O}$
14. $4 \text{Sm}_3\text{Zr}_4\text{P}_6\text{O}_{24} = 4 \text{SmZr}_2\text{P}_3\text{O}_{12} + 8 \text{SmPO}_4 + 8 \text{ZrO}_2 + \text{P}_4$
15. $7 \text{H}_2\text{SO}_4 + \text{C}_2\text{H}_6 = 2 \text{CO}_2 + 7 \text{SO}_2 + 10 \text{H}_2\text{O}$
16. $\text{FeS} + 2 \text{HCl} = \text{FeCl}_2 + \text{H}_2\text{S}$
17. $4 \text{Na}_2\text{S}_2\text{O}_3 + 4 \text{HCl} = 4 \text{NaCl} + 5 \text{SO}_2 + 3 \text{S} + 2 \text{H}_2\text{O}$
18. $12 \text{HClO}_4 + \text{P}_4\text{O}_{10} = 4 \text{H}_3\text{PO}_4 + 6 \text{Cl}_2\text{O}_7$
19. $\text{S}^{2-} + 2 \text{H}^{+} + \text{H}_2\text{O}_2 = \text{S} + 2 \text{H}_2\text{O}$



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