



BALANCE THE GIVEN CHEMICAL EQUATIONS

Worksheet - 34

- $\text{Mg}_3\text{N}_2 + \text{H}_2\text{O} = 3 \text{Mg}(\text{OH})_2 + \text{NH}_3$
- $\text{H}_2\text{S} + \text{Cu}(\text{NO}_3)_2 = \text{HNO}_3 + \text{SCu}$
- $\text{C}_7\text{H}_{16} + \text{O}_2 = \text{CO}_2 + 8 \text{H}_2\text{O}$
- $\text{WO}_3(\text{s}) + \text{H}_2(\text{g}) = \text{W}(\text{s}) + \text{H}_2\text{O}(\text{l})$
- $\text{K}_2\text{SO}_4 + 2 \text{NaCl} = \text{KCl} + \text{Na}_2\text{SO}_4$
- $\text{Mg}_3\text{N}_2 + \text{H}_2\text{O} = 3 \text{Mg}(\text{OH})_2 + 2 \text{NH}_3$
- $\text{SiF}_4 + 2 \text{H}_2\text{O} = \text{HF} + \text{SiO}_2$
- $4 \text{C}_2\text{H}_5\text{NH}_2(\text{g}) + \text{O}_2(\text{g}) = 8 \text{CO}_2(\text{g}) + 14 \text{H}_2\text{O}(\text{g}) + \text{N}_2(\text{g})$
- $\text{B}_2\text{Br}_6 + \text{HNO}_3 = 2 \text{B}(\text{NO}_3)_3 + \text{HBr}$
- $6 \text{NO} + \text{NH}_3 = 5 \text{N}_2 + \text{H}_2\text{O}$
- $6 \text{Li} + \text{N}_2 = \text{Li}_3\text{N}$
- $\text{K}_2\text{PtCl}_4 + \text{NH}_3 = \text{Pt}(\text{NH}_3)_2\text{Cl}_2 + \text{KCl}$
- $10 \text{S} + 40 \text{HNO}_3 = \text{H}_2\text{SO}_4 + 40 \text{NO}_2 + \text{H}_2\text{O}$
- $\text{KI} + \text{Pb}(\text{NO}_3)_2 = \text{PbI}_2 + 2 \text{KNO}_3$
- $3 \text{Cl}_2 + \text{NaOH} = 5 \text{NaCl} + \text{NaClO}_3 + 3 \text{H}_2\text{O}$
- $\text{Li}_2\text{CO}_3 + 2 \text{HNO}_3 = \text{LiNO}_3 + \text{H}_2\text{CO}_3$
- $\text{KAl}(\text{OH})_4(\text{aq}) + \text{H}_2\text{SO}_4(\text{aq}) + \text{H}_2\text{O}(\text{l}) = \text{KAl}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}(\text{s})$
- $57 \text{H}_2\text{S}_2\text{O}_3 + \text{H}_3\text{AsO}_3 = 8 \text{AsS}_3 + \text{SO}_2\text{H}_2\text{S}_4\text{O}_6 + 51 \text{H}_2\text{O}$
- $4 \text{Rb} + \text{O}_2 = \text{Rb}_2\text{O}$
- $2 \text{Ca}_3(\text{PO}_4)_2 + \text{SiO}_2 + 10 \text{C} = \text{CaSiO}_3 + \text{P}_4 + 10 \text{CO}$



ANSWERS

1. $\text{Mg}_3\text{N}_2 + 6 \text{H}_2\text{O} = 3 \text{Mg}(\text{OH})_2 + 2 \text{NH}_3$
2. $\text{H}_2\text{S} + \text{Cu}(\text{NO}_3)_2 = 2 \text{HNO}_3 + \text{SCu}$
3. $\text{C}_7\text{H}_{16} + 11 \text{O}_2 = 7 \text{CO}_2 + 8 \text{H}_2\text{O}$
4. $\text{WO}_3(\text{s}) + 3 \text{H}_2(\text{g}) = \text{W}(\text{s}) + 3 \text{H}_2\text{O}(\text{l})$
5. $\text{K}_2\text{SO}_4 + 2 \text{NaCl} = 2 \text{KCl} + \text{Na}_2\text{SO}_4$
6. $\text{Mg}_3\text{N}_2 + 6 \text{H}_2\text{O} = 3 \text{Mg}(\text{OH})_2 + 2 \text{NH}_3$
7. $\text{SiF}_4 + 2 \text{H}_2\text{O} = 4 \text{HF} + \text{SiO}_2$
8. $4 \text{C}_2\text{H}_5\text{NH}_2(\text{g}) + 15 \text{O}_2(\text{g}) = 8 \text{CO}_2(\text{g}) + 14 \text{H}_2\text{O}(\text{g}) + 2 \text{N}_2(\text{g})$
9. $\text{B}_2\text{Br}_6 + 6 \text{HNO}_3 = 2 \text{B}(\text{NO}_3)_3 + 6 \text{HBr}$
10. $6 \text{NO} + 4 \text{NH}_3 = 5 \text{N}_2 + 6 \text{H}_2\text{O}$
11. $6 \text{Li} + \text{N}_2 = 2 \text{Li}_3\text{N}$
12. $\text{K}_2\text{PtCl}_4 + 2 \text{NH}_3 = \text{Pt}(\text{NH}_3)_2\text{Cl}_2 + 2 \text{KCl}$
13. $10 \text{S} + 40 \text{HNO}_3 = 10 \text{H}_2\text{SO}_4 + 40 \text{NO}_2 + \text{H}_2\text{O}$
14. $2 \text{KI} + \text{Pb}(\text{NO}_3)_2 = \text{PbI}_2 + 2 \text{KNO}_3$
15. $3 \text{Cl}_2 + 6 \text{NaOH} = 5 \text{NaCl} + \text{NaClO}_3 + 3 \text{H}_2\text{O}$
16. $\text{Li}_2\text{CO}_3 + 2 \text{HNO}_3 = 2 \text{LiNO}_3 + \text{H}_2\text{CO}_3$
17. $\text{KAl}(\text{OH})_4(\text{aq}) + 2 \text{H}_2\text{SO}_4(\text{aq}) + 8 \text{H}_2\text{O}(\text{l}) = \text{KAl}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}(\text{s})$
18. $57 \text{H}_2\text{S}_2\text{O}_3 + 8 \text{H}_3\text{AsO}_3 = 8 \text{AsS}_3 + 18 \text{SO}_2 + \text{H}_2\text{S}_4\text{O}_6 + 51 \text{H}_2\text{O}$
19. $4 \text{Rb} + \text{O}_2 = 2 \text{Rb}_2\text{O}$
20. $2 \text{Ca}_3(\text{PO}_4)_2 + 6 \text{SiO}_2 + 10 \text{C} = 6 \text{CaSiO}_3 + \text{P}_4 + 10 \text{CO}$



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