



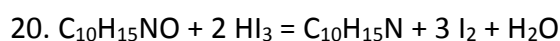
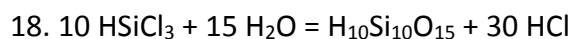
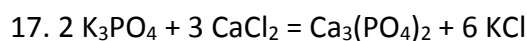
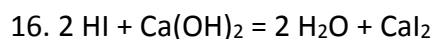
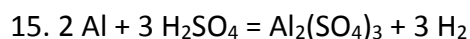
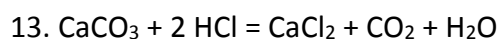
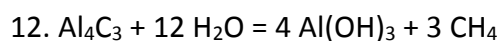
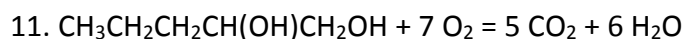
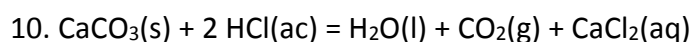
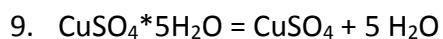
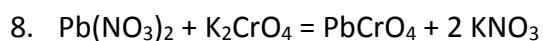
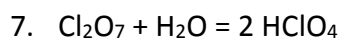
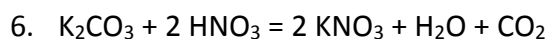
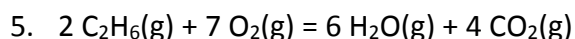
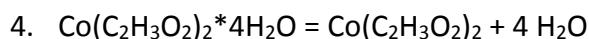
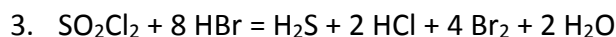
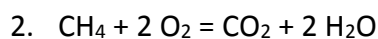
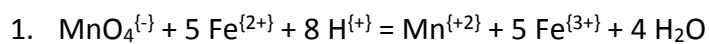
BALANCE THE GIVEN CHEMICAL EQUATIONS

Worksheet - 40

- $\text{MnO}_4^{-} + \text{Fe}^{2+} + 8 \text{H}^{+} = \text{Mn}^{2+} + \text{Fe}^{3+} + 4 \text{H}_2\text{O}$
- $\text{CH}_4 + 2 \text{O}_2 = \text{CO}_2 + \text{H}_2\text{O}$
- $\text{SO}_2\text{Cl}_2 + \text{HBr} = \text{H}_2\text{S} + \text{HCl} + 4 \text{Br}_2 + 2 \text{H}_2\text{O}$
- $\text{Co}(\text{C}_2\text{H}_3\text{O}_2)_2 \cdot 4\text{H}_2\text{O} = \text{Co}(\text{C}_2\text{H}_3\text{O}_2)_2 + \text{H}_2\text{O}$
- $2 \text{C}_2\text{H}_6(\text{g}) + \text{O}_2(\text{g}) = \text{H}_2\text{O}(\text{g}) + 4 \text{CO}_2(\text{g})$
- $\text{K}_2\text{CO}_3 + \text{HNO}_3 = \text{KNO}_3 + \text{H}_2\text{O} + \text{CO}_2$
- $\text{Cl}_2\text{O}_7 + \text{H}_2\text{O} = \text{HClO}_4$
- $\text{Pb}(\text{NO}_3)_2 + \text{K}_2\text{CrO}_4 = \text{PbCrO}_4 + \text{KNO}_3$
- $\text{CuSO}_4 \cdot 5\text{H}_2\text{O} = \text{CuSO}_4 + \text{H}_2\text{O}$
- $\text{CaCO}_3(\text{s}) + \text{HCl}(\text{ac}) = \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g}) + \text{CaCl}_2(\text{aq})$
- $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}(\text{OH})\text{CH}_2\text{OH} + \text{O}_2 = 5 \text{CO}_2 + \text{H}_2\text{O}$
- $\text{Al}_4\text{C}_3 + 12 \text{H}_2\text{O} = \text{Al}(\text{OH})_3 + 3 \text{CH}_4$
- $\text{CaCO}_3 + \text{HCl} = \text{CaCl}_2 + \text{CO}_2 + \text{H}_2\text{O}$
- $\text{Sr}(\text{NO}_3)_2 + \text{Li}_2\text{SO}_4 = \text{SrSO}_4 + 2 \text{LiNO}_3$
- $2 \text{Al} + \text{H}_2\text{SO}_4 = \text{Al}_2(\text{SO}_4)_3 + 3 \text{H}_2$
- $\text{HI} + \text{Ca}(\text{OH})_2 = \text{H}_2\text{O} + \text{CaI}_2$
- $2 \text{K}_3\text{PO}_4 + \text{CaCl}_2 = \text{Ca}_3(\text{PO}_4)_2 + 6 \text{KCl}$
- $10 \text{HSiCl}_3 + \text{H}_2\text{O} = \text{H}_{10}\text{Si}_{10}\text{O}_{15} + 30 \text{HCl}$
- $\text{NaOH} + \text{NaNO}_2 + \text{Al} + \text{H}_2\text{O} = \text{NH}_3 + \text{NaAlO}_2$
- $\text{C}_{10}\text{H}_{15}\text{NO} + \text{HI}_3 = \text{C}_{10}\text{H}_{15}\text{N} + \text{I}_2 + \text{H}_2\text{O}$



ANSWERS





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