



BALANCE THE GIVEN CHEMICAL EQUATIONS

Worksheet - 41

- $\text{TiCl}_4 + \text{H}_2\text{O} = \text{TiO}_2 + \text{HCl}$
- $\text{K}_2\text{SO}_3 + \text{Mn}(\text{OH})_2 = \text{KOH} + \text{MnSO}_3$
- $\text{Ni} + \text{HCl} = \text{NiCl}_2 + \text{H}_2$
- $\text{Cr}_2\text{O}_3 + \text{Cl}_2 + 3 \text{C} = \text{CrCl}_3 + 3 \text{CO}$
- $2 \text{C}_2\text{H}_2 + \text{O}_2 = 4 \text{CO}_2 + \text{H}_2\text{O}$
- $2 \text{CH}_3\text{OH}(\text{l}) + \text{O}_2(\text{g}) = 2 \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$
- $2 \text{C}_3\text{H}_8\text{O} + \text{O}_2 = 6 \text{CO}_2 + \text{H}_2\text{O}$
- $2 \text{FeS}_2 + \text{O}_2 = \text{Fe}_2\text{O}_2 + \text{SO}_2$
- $\text{Sn} + \text{NaOH} = \text{Na}_2\text{SnO}_2 + \text{H}_2$
- $\text{Pb}(\text{NO}_3)_2 + \text{NaCl} = \text{PbCl}_2 + \text{NaNO}_3$
- $2 \text{HCl} + \text{Ba}(\text{OH})_2 = \text{BaCl}_2 + \text{H}_2\text{O}$
- $\text{CaO} + \text{C} = \text{CaC}_2 + \text{CO}$
- $3 \text{Mg}(\text{OH})_2 + \text{H}_3\text{PO}_4 = \text{Mg}_3(\text{PO}_4)_2 + \text{HOH}$
- $\text{MnO}_2 + \text{HCl} = \text{MnCl}_2 + \text{Cl}_2 + \text{H}_2\text{O}$
- $\text{UO}_2(\text{s}) + \text{HF}(\text{l}) = \text{UF}_4(\text{s}) + \text{H}_2\text{O}(\text{l})$
- $\text{Cr}_2\text{O}_7^{2-} + \text{TeO}_2 + 8 \text{H}^{(+)} = \text{Cr}^{3+} + 3 \text{H}_2\text{TeO}_4 + \text{H}_2\text{O}$
- $3 \text{P} + \text{HNO}_3 + 2 \text{H}_2\text{O} = \text{H}_3\text{PO}_4 + 5 \text{NO}$
- $2 \text{Al} + \text{Cl}_2 = \text{AlCl}_3$
- $3 \text{NO}_2 + \text{H}_2\text{O} = \text{HNO}_3 + \text{NO}$
- $\text{NH}_4\text{ClO}_4 = \text{N}_2 + \text{Cl}_2 + \text{O}_2 + 4 \text{H}_2\text{O}$



ANSWERS

1. $\text{TiCl}_4 + 2 \text{H}_2\text{O} = \text{TiO}_2 + 4 \text{HCl}$
2. $\text{K}_2\text{SO}_3 + \text{Mn}(\text{OH})_2 = 2 \text{KOH} + \text{MnSO}_3$
3. $\text{Ni} + 2 \text{HCl} = \text{NiCl}_2 + \text{H}_2$
4. $\text{Cr}_2\text{O}_3 + 3 \text{Cl}_2 + 3 \text{C} = 2 \text{CrCl}_3 + 3 \text{CO}$
5. $2 \text{C}_2\text{H}_2 + 5 \text{O}_2 = 4 \text{CO}_2 + 2 \text{H}_2\text{O}$
6. $2 \text{CH}_3\text{OH}(\text{l}) + 3 \text{O}_2(\text{g}) = 2 \text{CO}_2(\text{g}) + 4 \text{H}_2\text{O}(\text{g})$
7. $2 \text{C}_3\text{H}_8\text{O} + 9 \text{O}_2 = 6 \text{CO}_2 + 8 \text{H}_2\text{O}$
8. $2 \text{FeS}_2 + 5 \text{O}_2 = \text{Fe}_2\text{O}_3 + 4 \text{SO}_2$
9. $\text{Sn} + 2 \text{NaOH} = \text{Na}_2\text{SnO}_2 + \text{H}_2$
10. $\text{Pb}(\text{NO}_3)_2 + 2 \text{NaCl} = \text{PbCl}_2 + 2 \text{NaNO}_3$
11. $2 \text{HCl} + \text{Ba}(\text{OH})_2 = \text{BaCl}_2 + 2 \text{H}_2\text{O}$
12. $\text{CaO} + 3 \text{C} = \text{CaC}_2 + \text{CO}$
13. $3 \text{Mg}(\text{OH})_2 + 2 \text{H}_3\text{PO}_4 = \text{Mg}_3(\text{PO}_4)_2 + 6 \text{HOH}$
14. $\text{MnO}_2 + 4 \text{HCl} = \text{MnCl}_2 + \text{Cl}_2 + 2 \text{H}_2\text{O}$
15. $\text{UO}_2(\text{s}) + 4 \text{HF}(\text{l}) = \text{UF}_4(\text{s}) + 2 \text{H}_2\text{O}(\text{l})$
16. $\text{Cr}_2\text{O}_7^{2-} + 3 \text{TeO}_2 + 8 \text{H}^{(+)} = 2 \text{Cr}^{3+} + 3 \text{H}_2\text{TeO}_4 + \text{H}_2\text{O}$
17. $3 \text{P} + 5 \text{HNO}_3 + 2 \text{H}_2\text{O} = 3 \text{H}_3\text{PO}_4 + 5 \text{NO}$
18. $2 \text{Al} + 3 \text{Cl}_2 = 2 \text{AlCl}_3$
19. $3 \text{NO}_2 + \text{H}_2\text{O} = 2 \text{HNO}_3 + \text{NO}$
20. $2 \text{NH}_4\text{ClO}_4 = \text{N}_2 + \text{Cl}_2 + 2 \text{O}_2 + 4 \text{H}_2\text{O}$



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