



BALANCE THE GIVEN CHEMICAL EQUATIONS

Worksheet - 46

1. $2 (\text{CH}_3(\text{CH}_2)_4\text{CH}_3) + \text{ ______ } \text{O}_2 = \text{ ______ } \text{CO}_2 + 14 \text{H}_2\text{O}$
2. $\text{C}_3\text{H}_8 + \text{ ______ } \text{O}_2 = 3 \text{CO}_2 + \text{ ______ } \text{H}_2\text{O}$
3. $3 \text{Ca}(\text{OH})_2 + \text{ ______ } \text{H}_3\text{PO}_4 = \text{Ca}_3(\text{PO}_4)_2 + \text{ ______ } \text{H}_2\text{O}$
4. $\text{ ______ } \text{KClO}_3 = \text{ ______ } \text{KCl} + 3 \text{O}_2$
5. $\text{ ______ } \text{HCl} + \text{O}_2 = \text{ ______ } \text{Cl}_2 + 2 \text{H}_2\text{O}$
6. $\text{Ca}_3\text{P}_2 + \text{ ______ } \text{H}_2\text{O} = \text{ ______ } \text{Ca}(\text{OH})_2 + 2 \text{PH}_3$
7. $20 \text{LiOH} + \text{ ______ } \text{CO}_2 = \text{ ______ } \text{LiCO}_3 + \text{H}_2\text{O}$
8. $\text{ ______ } \text{HClO}_4 + \text{Ba}(\text{OH})_2 = \text{Ba}(\text{ClO}_4)_2 + \text{ ______ } \text{H}_2\text{O}$
9. $\text{ ______ } \text{FeCl}_3 + \text{H}_2\text{S} = \text{ ______ } \text{FeCl}_2 + 2 \text{HCl} + \text{S}$
10. $\text{Mg}_3\text{N}_2(\text{s}) + \text{ ______ } \text{H}_2\text{O}(\text{l}) = 3 \text{Mg}(\text{OH})_2(\text{s}) + \text{ ______ } \text{NH}_3(\text{g})$
11. $\text{ ______ } \text{C}_4\text{H}_{10} + 13 \text{O}_2 = \text{ ______ } \text{CO}_2 + 10 \text{H}_2\text{O}$
12. $\text{Fe}_2\text{O}_3 + \text{ ______ } \text{H}_2 = \text{ ______ } \text{H}_2\text{O} + 2 \text{Fe}$
13. $2 \text{C}_2\text{H}_2 + \text{ ______ } \text{O}_2 = \text{ ______ } \text{CO}_2 + 2 \text{H}_2\text{O}$
14. $\text{N}_2 + \text{ ______ } \text{H}_2 = \text{ ______ } \text{NH}_3$
15. $2 \text{C}_8\text{H}_{18} + \text{ ______ } \text{O}_2 = \text{ ______ } \text{CO}_2 + 18 \text{H}_2\text{O}$
16. $2 \text{C}_3\text{H}_8\text{O} + \text{ ______ } \text{O}_2 = \text{ ______ } \text{CO}_2 + 8 \text{H}_2\text{O}$
17. $\text{Na}_2\text{SO}_3 + \text{ ______ } \text{HNO}_3 = \text{ ______ } \text{NaNO}_3 + \text{SO}_2 + \text{H}_2\text{O}$
18. $2 (\text{CH}_3\text{CH}_3) + \text{ ______ } \text{O}_2 = \text{ ______ } \text{CO}_2 + 6 \text{H}_2\text{O}$
19. $2 \text{H}_2\text{S} + \text{ ______ } \text{O}_2 = \text{ ______ } \text{H}_2\text{O} + 2 \text{SO}_2$
20. $\text{B}_2\text{Br}_6 + \text{ ______ } \text{HNO}_3 = 2 \text{B}(\text{NO}_3)_3 + \text{ ______ } \text{HBr}$



ANSWERS

1. $2 (\text{CH}_3(\text{CH}_2)_4\text{CH}_3) + 19 \text{O}_2 = 12 \text{CO}_2 + 14 \text{H}_2\text{O}$
2. $\text{C}_3\text{H}_8 + 5 \text{O}_2 = 3 \text{CO}_2 + 4 \text{H}_2\text{O}$
3. $3 \text{Ca}(\text{OH})_2 + 2 \text{H}_3\text{PO}_4 = \text{Ca}_3(\text{PO}_4)_2 + 6 \text{H}_2\text{O}$
4. $2 \text{KClO}_3 = 2 \text{KCl} + 3 \text{O}_2$
5. $4 \text{HCl} + \text{O}_2 = 2 \text{Cl}_2 + 2 \text{H}_2\text{O}$
6. $\text{Ca}_3\text{P}_2 + 6 \text{H}_2\text{O} = 3 \text{Ca}(\text{OH})_2 + 2 \text{PH}_3$
7. $20 \text{LiOH} + 20 \text{CO}_2 = 20 \text{LiCO}_3 + \text{H}_2\text{O}$
8. $2 \text{HClO}_4 + \text{Ba}(\text{OH})_2 = \text{Ba}(\text{ClO}_4)_2 + 2 \text{H}_2\text{O}$
9. $2 \text{FeCl}_3 + \text{H}_2\text{S} = 2 \text{FeCl}_2 + 2 \text{HCl} + \text{S}$
10. $\text{Mg}_3\text{N}_2(\text{s}) + 6 \text{H}_2\text{O}(\text{l}) = 3 \text{Mg}(\text{OH})_2(\text{s}) + 2 \text{NH}_3(\text{g})$
11. $2 \text{C}_4\text{H}_{10} + 13 \text{O}_2 = 8 \text{CO}_2 + 10 \text{H}_2\text{O}$
12. $\text{Fe}_2\text{O}_3 + 3 \text{H}_2 = 3 \text{H}_2\text{O} + 2 \text{Fe}$
13. $2 \text{C}_2\text{H}_2 + 5 \text{O}_2 = 4 \text{CO}_2 + 2 \text{H}_2\text{O}$
14. $\text{N}_2 + 3 \text{H}_2 = 2 \text{NH}_3$
15. $2 \text{C}_8\text{H}_{18} + 25 \text{O}_2 = 16 \text{CO}_2 + 18 \text{H}_2\text{O}$
16. $2 \text{C}_3\text{H}_8\text{O} + 9 \text{O}_2 = 6 \text{CO}_2 + 8 \text{H}_2\text{O}$
17. $\text{Na}_2\text{SO}_3 + 2 \text{HNO}_3 = 2 \text{NaNO}_3 + \text{SO}_2 + \text{H}_2\text{O}$
18. $2 (\text{CH}_3\text{CH}_3) + 7 \text{O}_2 = 4 \text{CO}_2 + 6 \text{H}_2\text{O}$
19. $2 \text{H}_2\text{S} + 3 \text{O}_2 = 2 \text{H}_2\text{O} + 2 \text{SO}_2$
20. $\text{B}_2\text{Br}_6 + 6 \text{HNO}_3 = 2 \text{B}(\text{NO}_3)_3 + 6 \text{HBr}$



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