



BALANCE THE GIVEN CHEMICAL EQUATIONS

Worksheet - 51

- $4 \text{P}_3 + \text{O}_2 = \text{P}_2\text{O}_3$
- $\text{CH}_3\text{CH}_3 + 7 \text{O}_2 = \text{CO}_2 + 6 \text{H}_2\text{O}$
- $\text{TiCl}_4 + \text{H}_2\text{O} = \text{TiO}_2 + \text{HCl}$
- $2 \text{GaBr}_3 + \text{Cl}_2 = \text{GaCl}_3 + 3 \text{Br}_2$
- $\text{H}_2\text{C}_2\text{O}_4 + \text{OH} = \text{H}_2\text{O} + \text{C}_2\text{O}_4$
- $\text{CO} + 17 \text{H}_2 = \text{C}_8\text{H}_{18} + \text{H}_2\text{O}$
- $\text{KMnO}_4 + 16 \text{HCl} = \text{Cl}_2 + 2 \text{MnCl}_2 + 2 \text{KCl} + 8 \text{H}_2\text{O}$
- $\text{Fe} + 3 \text{O}_2 = \text{Fe}_2\text{O}_3$
- $2 \text{K} + \text{H}_2\text{O} = \text{KOH} + \text{H}_2$
- $2 \text{C}_4\text{H}_6 + \text{O}_2 = \text{CO}_2 + 6 \text{H}_2\text{O}$
- $\text{Na}_3\text{PO}_4 + 3 \text{NiCl}_2 = \text{NaCl} + \text{Ni}_3(\text{PO}_4)_2$
- $\text{Pb}(\text{SO}_4) + \text{H}_2\text{O} = \text{PbO}_2 + \text{SO}_4 + \text{H}_3\text{O}$
- $\text{Al} + 3 \text{S}_8 = \text{Al}_2\text{S}_3$
- $10 \text{C}_3\text{H}_8 + \text{O}_2 = \text{CO} + 4 \text{H}_2\text{O}$
- $\text{Pb} + \text{O}_2 + 2 \text{H}_2\text{O} = \text{Pb}(\text{OH})_2$
- $2 \text{NBr}_3 + \text{NaOH} = \text{N}_2 + \text{NaBr} + 3 \text{HOBr}$
- $\text{C}_3\text{H}_8 + \text{O}_2 = \text{CO}_2 + 4 \text{H}_2\text{O}$
- $\text{Hg} + \text{H}_2\text{SO}_4 = \text{HgSO}_4 + \text{H}_2\text{O} + \text{SO}_2$
- $2 \text{C}_3\text{H}_8\text{O} + \text{O}_2 = \text{CO}_2 + 8 \text{H}_2\text{O}$
- $\text{HClO}_4 + \text{P}_4\text{O}_{10} = \text{H}_3\text{PO}_4 + 6 \text{Cl}_2\text{O}_7$



ANSWERS

1. $4 \text{P}_3 + 9 \text{O}_2 = 6 \text{P}_2\text{O}_3$
2. $2 \text{CH}_3\text{CH}_3 + 7 \text{O}_2 = 4 \text{CO}_2 + 6 \text{H}_2\text{O}$
3. $\text{TiCl}_4 + 2 \text{H}_2\text{O} = \text{TiO}_2 + 4 \text{HCl}$
4. $2 \text{GaBr}_3 + 3 \text{Cl}_2 = 2 \text{GaCl}_3 + 3 \text{Br}_2$
5. $\text{H}_2\text{C}_2\text{O}_4 + 2 \text{OH} = 2 \text{H}_2\text{O} + \text{C}_2\text{O}_4$
6. $8 \text{CO} + 17 \text{H}_2 = \text{C}_8\text{H}_{18} + 8 \text{H}_2\text{O}$
7. $2 \text{KMnO}_4 + 16 \text{HCl} = 5 \text{Cl}_2 + 2 \text{MnCl}_2 + 2 \text{KCl} + 8 \text{H}_2\text{O}$
8. $4 \text{Fe} + 3 \text{O}_2 = 2 \text{Fe}_2\text{O}_3$
9. $2 \text{K} + 2 \text{H}_2\text{O} = 2 \text{KOH} + \text{H}_2$
10. $2 \text{C}_4\text{H}_6 + 11 \text{O}_2 = 8 \text{CO}_2 + 6 \text{H}_2\text{O}$
11. $2 \text{Na}_3\text{PO}_4 + 3 \text{NiCl}_2 = 6 \text{NaCl} + \text{Ni}_3(\text{PO}_4)_2$
12. $\text{Pb}(\text{SO}_4) + 6 \text{H}_2\text{O} = \text{PbO}_2 + \text{SO}_4 + 4 \text{H}_3\text{O}$
13. $16 \text{Al} + 3 \text{S}_8 = 8 \text{Al}_2\text{S}_3$
14. $10 \text{C}_3\text{H}_8 + 15 \text{O}_2 = 30 \text{CO} + 4 \text{H}_{20}$
15. $2 \text{Pb} + \text{O}_2 + 2 \text{H}_2\text{O} = 2 \text{Pb}(\text{OH})_2$
16. $2 \text{NBr}_3 + 3 \text{NaOH} = \text{N}_2 + 3 \text{NaBr} + 3 \text{HOBr}$
17. $\text{C}_3\text{H}_8 + 5 \text{O}_2 = 3 \text{CO}_2 + 4 \text{H}_2\text{O}$
18. $\text{Hg} + 2 \text{H}_2\text{SO}_4 = \text{HgSO}_4 + 2 \text{H}_2\text{O} + \text{SO}_2$
19. $2 \text{C}_3\text{H}_8\text{O} + 9 \text{O}_2 = 6 \text{CO}_2 + 8 \text{H}_2\text{O}$
20. $12 \text{HClO}_4 + \text{P}_4\text{O}_{10} = 4 \text{H}_3\text{PO}_4 + 6 \text{Cl}_2\text{O}_7$



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