



BALANCE THE GIVEN CHEMICAL EQUATIONS

Worksheet - 55

- $P_2O_5 + \text{ ____ } H_2O = \text{ ____ } H_3PO_4$
- $\text{ ____ } CO + I_2O_5 = \text{ ____ } CO_2 + I_2$
- $2 C_8H_{18} + \text{ ____ } O_2 = \text{ ____ } CO_2 + 18 H_2O$
- $MnO_2 + \text{ ____ } HCl = MnCl_2 + \text{ ____ } H_2O + Cl_2$
- $MnO_2 + 2 Cl^{-} + \text{ ____ } H^{+} = Cl_2 + Mn^{2+} + \text{ ____ } H_2O$
- $\text{ ____ } CH_3(CH_2)_2CH_3 + 13 O_2 = \text{ ____ } CO_2 + 10 H_2O$
- $Fe_2O_3 + \text{ ____ } CO = \text{ ____ } Fe + 3 CO_2$
- $2 Eu + \text{ ____ } HF = 2 EuF_3 + \text{ ____ } H_2$
- $\text{ ____ } KCrO_2 + 3 Br_2 + 8 KOH = \text{ ____ } K_2CrO_4 + 6 KBr + 4 H_2O$
- $Na_2 + \text{ ____ } H_2O = \text{ ____ } NaOH + H_2$
- $\text{ ____ } NaClO_2 = NaCl + \text{ ____ } NaClO_3$
- $\text{ ____ } P + 5 O_2 = \text{ ____ } P_2O_5$
- $3 Na_2CO_3 + \text{ ____ } H_3PO_4 = 2 Na_3PO_4 + \text{ ____ } H_2O + 3 CO_2$
- $\text{ ____ } Al + 3 O_2 = \text{ ____ } Al_2O_3$
- $2 Al + \text{ ____ } HCl = 2 AlCl_3 + \text{ ____ } H_2$
- $C_3H_8 + \text{ ____ } O_2 = \text{ ____ } CO_2 + 4 H_2O$
- $2 C_2H_6 + \text{ ____ } O_2 = 4 CO_2 + \text{ ____ } H_2O$
- $\text{ ____ } N_2O_5 = \text{ ____ } NO_2 + O_2$
- $H_2SiO_3 + \text{ ____ } KOH = K_2SiO_3 + \text{ ____ } H_2O$
- $Al(OH)_3 + \text{ ____ } HCl = AlCl_3 + \text{ ____ } H_2O$



ANSWERS

1. $\text{P}_2\text{O}_5 + 3 \text{H}_2\text{O} = 2 \text{H}_3\text{PO}_4$
2. $5 \text{CO} + \text{I}_2\text{O}_5 = 5 \text{CO}_2 + \text{I}_2$
3. $2 \text{C}_8\text{H}_{18} + 25 \text{O}_2 = 16 \text{CO}_2 + 18 \text{H}_2\text{O}$
4. $\text{MnO}_2 + 4 \text{HCl} = \text{MnCl}_2 + 2 \text{H}_2\text{O} + \text{Cl}_2$
5. $\text{MnO}_2 + 2 \text{Cl}^{-} + 4 \text{H}^{+} = \text{Cl}_2 + \text{Mn}^{2+} + 2 \text{H}_2\text{O}$
6. $2 \text{CH}_3(\text{CH}_2)_2\text{CH}_3 + 13 \text{O}_2 = 8 \text{CO}_2 + 10 \text{H}_2\text{O}$
7. $\text{Fe}_2\text{O}_3 + 3 \text{CO} = 2 \text{Fe} + 3 \text{CO}_2$
8. $2 \text{Eu} + 6 \text{HF} = 2 \text{EuF}_3 + 3 \text{H}_2$
9. $2 \text{KCrO}_2 + 3 \text{Br}_2 + 8 \text{KOH} = 2 \text{K}_2\text{CrO}_4 + 6 \text{KBr} + 4 \text{H}_2\text{O}$
10. $\text{Na}_2 + 2 \text{H}_2\text{O} = 2 \text{NaOH} + \text{H}_2$
11. $3 \text{NaClO}_2 = \text{NaCl} + 2 \text{NaClO}_3$
12. $4 \text{P} + 5 \text{O}_2 = 2 \text{P}_2\text{O}_5$
13. $3 \text{Na}_2\text{CO}_3 + 2 \text{H}_3\text{PO}_4 = 2 \text{Na}_3\text{PO}_4 + 3 \text{H}_2\text{O} + 3 \text{CO}_2$
14. $4 \text{Al} + 3 \text{O}_2 = 2 \text{Al}_2\text{O}_3$
15. $2 \text{Al} + 6 \text{HCl} = 2 \text{AlCl}_3 + 3 \text{H}_2$
16. $\text{C}_3\text{H}_8 + 5 \text{O}_2 = 3 \text{CO}_2 + 4 \text{H}_2\text{O}$
17. $2 \text{C}_2\text{H}_6 + 7 \text{O}_2 = 4 \text{CO}_2 + 6 \text{H}_2\text{O}$
18. $2 \text{N}_2\text{O}_5 = 4 \text{NO}_2 + \text{O}_2$
19. $\text{H}_2\text{SiO}_3 + 2 \text{KOH} = \text{K}_2\text{SiO}_3 + 2 \text{H}_2\text{O}$
20. $\text{Al}(\text{OH})_3 + 3 \text{HCl} = \text{AlCl}_3 + 3 \text{H}_2\text{O}$



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