



## BALANCE THE GIVEN CHEMICAL EQUATIONS

### Worksheet - 6

- $4 \text{NH}_3 + \text{O}_2 = 4 \text{NO} + \text{H}_2\text{O}$
- $2 \text{Fe(s)} + \text{O(g)} = \text{Fe}_2\text{O}_3\text{(s)}$
- $\text{C}_2\text{H}_6\text{O} + 5 \text{O}_2 = \text{H}_2\text{O} + 4 \text{C}_2\text{H}_3\text{O}_2$
- $\text{Fe} + \text{O}_2 + 2 \text{H}_2\text{O} = \text{Fe(OH)}_2$
- $\text{LiHCO}_3 + \text{H}_2\text{SO}_4 = \text{Li}_2\text{SO}_4 + \text{H}_2\text{O} + 2 \text{CO}_2$
- $\text{C}_3\text{H}_8 + \text{O}_2 = 3 \text{CO}_2 + 4 \text{H}_2\text{O}$
- $2 \text{C}_4\text{H}_{10} + \text{O}_2 = 8 \text{CO}_2 + \text{H}_2\text{O}$
- $2 \text{C}_4\text{H}_{10} + \text{O}_2 = 8 \text{CO}_2 + \text{H}_2\text{O}$
- $\text{N}_2 + \text{H}_2 = 2 \text{NH}_3$
- $\text{Al(OH)}_3 + \text{HCl} = \text{AlCl}_3 + \text{H}_2\text{O}$
- $8 \text{Al} + \text{Fe}_3\text{O}_4 = 4 \text{Al}_2\text{O}_3 + 9 \text{Fe}$
- $2 \text{KBr} + \text{H}_2\text{SO}_4 = \text{K}_2\text{SO}_4 + \text{Br}_2 + \text{SO}_2 + 2 \text{H}_2\text{O}$
- $3 \text{MgO}_2 + 4 \text{Al} = 2 \text{Al}_2\text{O}_3 + \text{Mg}$
- $2 \text{C}_3\text{H}_8\text{O} + \text{O}_2 = 6 \text{CO}_2 + 8 \text{H}_2\text{O}$
- $2 \text{Ca} + \text{O}_2 = \text{CaO}$
- $\text{CH}_4 + 9 \text{O}_2 = \text{CO}_2 + 2 \text{H}_2\text{O} + 4 \text{SO}_2$
- $\text{Li} + \text{O}_2 = \text{Li}_2\text{O}$
- $2 \text{Al} + \text{Br}_2 = \text{Al}_2\text{Br}_6$
- $((\text{NH}_4)_3\text{PO}_4) = \text{H}_3\text{PO}_4 + \text{NH}_3$
- $\text{Na} + \text{FeCl}_3 = \text{Fe} + \text{NaCl}$



# ANSWERS

1.  $4 \text{NH}_3 + 5 \text{O}_2 = 4 \text{NO} + 6 \text{H}_2\text{O}$
2.  $2 \text{Fe(s)} + 3 \text{O(g)} = \text{Fe}_2\text{O}_3\text{(s)}$
3.  $4 \text{C}_2\text{H}_6\text{O} + 5 \text{O}_2 = 6 \text{H}_2\text{O} + 4 \text{C}_2\text{H}_3\text{O}_2$
4.  $2 \text{Fe} + \text{O}_2 + 2 \text{H}_2\text{O} = 2 \text{Fe(OH)}_2$
5.  $2 \text{LiHCO}_3 + \text{H}_2\text{SO}_4 = \text{Li}_2\text{SO}_4 + 2 \text{H}_2\text{O} + 2 \text{CO}_2$
6.  $\text{C}_3\text{H}_8 + 5 \text{O}_2 = 3 \text{CO}_2 + 4 \text{H}_2\text{O}$
7.  $2 \text{C}_4\text{H}_{10} + 8 \text{O}_2 = 8 \text{CO}_2 + \text{H}_2\text{O}$
8.  $2 \text{C}_4\text{H}_{10} + 13 \text{O}_2 = 8 \text{CO}_2 + 10 \text{H}_2\text{O}$
9.  $\text{N}_2 + 3 \text{H}_2 = 2 \text{NH}_3$
10.  $\text{Al(OH)}_3 + 3 \text{HCl} = \text{AlCl}_3 + 3 \text{H}_2\text{O}$
11.  $8 \text{Al} + 3 \text{Fe}_3\text{O}_4 = 4 \text{Al}_2\text{O}_3 + 9 \text{Fe}$
12.  $2 \text{KBr} + 2 \text{H}_2\text{SO}_4 = \text{K}_2\text{SO}_4 + \text{Br}_2 + \text{SO}_2 + 2 \text{H}_2\text{O}$
13.  $3 \text{MgO}_2 + 4 \text{Al} = 2 \text{Al}_2\text{O}_3 + 3 \text{Mg}$
14.  $2 \text{C}_3\text{H}_8\text{O} + 9 \text{O}_2 = 6 \text{CO}_2 + 8 \text{H}_2\text{O}$
15.  $2 \text{Ca} + \text{O}_2 = 2 \text{CaO}$
16.  $4 \text{CHS} + 9 \text{O}_2 = 4 \text{CO}_2 + 2 \text{H}_2\text{O} + 4 \text{SO}_2$
17.  $4 \text{Li} + \text{O}_2 = 2 \text{Li}_2\text{O}$
18.  $2 \text{Al} + 3 \text{Br}_2 = \text{Al}_2\text{Br}_6$
19.  $((\text{NH}_4)_3\text{PO}_4) = \text{H}_3\text{PO}_4 + 3 \text{NH}_3$
20.  $3 \text{Na} + \text{FeCl}_3 = \text{Fe} + 3 \text{NaCl}$



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