



BALANCE THE GIVEN CHEMICAL EQUATIONS

Worksheet - 62

1. $\text{C}_6\text{H}_{14} + 19 \text{O}_2 = \text{CO}_2 + 14 \text{H}_2\text{O}$
2. $2 \text{C}_4\text{H}_8\text{O}(\text{g}) + \text{O}_2(\text{g}) = 8 \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$
3. $\text{PbS} + \text{H}_2\text{O}_2 = \text{PbSO}_4 + \text{H}_2\text{O}$
4. $\text{Al}_2\text{S}_3 + \text{H}_2\text{O} = 2 \text{Al}(\text{OH})_3 + \text{H}_2\text{S}$
5. $\text{FeS}_2 + 11 \text{O}_2 = \text{Fe}_2\text{O}_3 + 8 \text{SO}_2$
6. $2 \text{CH}_3(\text{CH}_2)_2\text{CH}_3 + \text{O}_2 = 8 \text{CO}_2 + \text{H}_2\text{O}$
7. $\text{Ag} + 4 \text{HNO}_3 = \text{NO} + \text{H}_2\text{O} + 3 \text{AgNO}_3$
8. $\text{C}_5\text{H}_{12} + \text{O}_2 = 5 \text{CO}_2 + \text{H}_2\text{O}$
9. $\text{S}_8(\text{s}) + \text{O}_2(\text{g}) = \text{SO}_2(\text{g})$
10. $\text{NH}_4\text{ClO}_4 = \text{N}_2 + \text{Cl}_2 + \text{O}_2 + 4 \text{H}_2\text{O}$
11. $\text{C}_5\text{H}_{12} + 11 \text{O}_2 = 10 \text{CO} + \text{H}_2\text{O}$
12. $\text{NH}_3 + 5 \text{O}_2 = 4 \text{NO} + \text{H}_2\text{O}$
13. $\text{Ag}_2\text{SO}_4 + \text{HNO}_3 = \text{AgNO}_3 + \text{H}_2\text{SO}_4$
14. $\text{SeO}_3^{\{-2\}} + \text{I}^{\{-\}} + 6 \text{H}^{\{+\}} = \text{Se} + 2 \text{I}_2 + \text{H}_2\text{O}$
15. $\text{CH}_4 + \text{O}_2 = \text{CO}_2 + \text{H}_2\text{O}$
16. $\text{Mn}_2\text{O}_3 + \text{H}_2\text{SO}_4 = \text{Mn}_2(\text{SO}_4)_3 + \text{H}_2\text{O}$
17. $\text{Pb}(\text{NO}_3)_2 = \text{PbO} + 4 \text{NO}_2 + \text{O}_2$
18. $6 \text{NaNO}_3 + \text{Al} = 3 \text{Na}_2\text{O} + \text{Al}_2\text{O}_3 + 3 \text{N}_2$
19. $\text{FeS} + 7 \text{O}_2 = 2 \text{Fe}_2\text{O}_3 + \text{SO}_2$
20. $20 \text{CO}_2 + \text{KOH} = \text{KCO}_3 + \text{H}_2\text{O}$



ANSWERS

1. $2 \text{C}_6\text{H}_{14} + 19 \text{O}_2 = 12 \text{CO}_2 + 14 \text{H}_2\text{O}$
2. $2 \text{C}_4\text{H}_8\text{O}(\text{g}) + 11 \text{O}_2(\text{g}) = 8 \text{CO}_2(\text{g}) + 8 \text{H}_2\text{O}(\text{g})$
3. $\text{PbS} + 4 \text{H}_2\text{O}_2 = \text{PbSO}_4 + 4 \text{H}_2\text{O}$
4. $\text{Al}_2\text{S}_3 + 6 \text{H}_2\text{O} = 2 \text{Al}(\text{OH})_3 + 3 \text{H}_2\text{S}$
5. $4 \text{FeS}_2 + 11 \text{O}_2 = 2 \text{Fe}_2\text{O}_3 + 8 \text{SO}_2$
6. $2 \text{CH}_3(\text{CH}_2)_2\text{CH}_3 + 13 \text{O}_2 = 8 \text{CO}_2 + 10 \text{H}_2\text{O}$
7. $3 \text{Ag} + 4 \text{HNO}_3 = \text{NO} + 2 \text{H}_2\text{O} + 3 \text{AgNO}_3$
8. $\text{C}_5\text{H}_{12} + 8 \text{O}_2 = 5 \text{CO}_2 + 6 \text{H}_2\text{O}$
9. $\text{S}_8(\text{s}) + 8 \text{O}_2(\text{g}) = 8 \text{SO}_2(\text{g})$
10. $2 \text{NH}_4\text{ClO}_4 = \text{N}_2 + \text{Cl}_2 + 2 \text{O}_2 + 4 \text{H}_2\text{O}$
11. $2 \text{C}_5\text{H}_{12} + 11 \text{O}_2 = 10 \text{CO} + 12 \text{H}_2\text{O}$
12. $4 \text{NH}_3 + 5 \text{O}_2 = 4 \text{NO} + 6 \text{H}_2\text{O}$
13. $\text{Ag}_2\text{SO}_4 + 2 \text{HNO}_3 = 2 \text{AgNO}_3 + \text{H}_2\text{SO}_4$
14. $\text{SeO}_3^{\{-2\}} + 4 \text{I}^{\{-\}} + 6 \text{H}^{\{+\}} = \text{Se} + 2 \text{I}_2 + 3 \text{H}_2\text{O}$
15. $\text{CH}_4 + 2 \text{O}_2 = \text{CO}_2 + 2 \text{H}_2\text{O}$
16. $\text{Mn}_2\text{O}_3 + 3 \text{H}_2\text{SO}_4 = \text{Mn}_2(\text{SO}_4)_3 + 3 \text{H}_2\text{O}$
17. $2 \text{Pb}(\text{NO}_3)_2 = 2 \text{PbO} + 4 \text{NO}_2 + \text{O}_2$
18. $6 \text{NaNO}_3 + 10 \text{Al} = 3 \text{Na}_2\text{O} + 5 \text{Al}_2\text{O}_3 + 3 \text{N}_2$
19. $4 \text{FeS} + 7 \text{O}_2 = 2 \text{Fe}_2\text{O}_3 + 4 \text{SO}_2$
20. $20 \text{CO}_2 + 20 \text{KOH} = 20 \text{KCO}_3 + \text{H}_2\text{O}$



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