



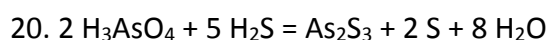
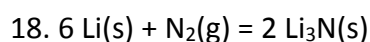
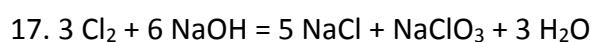
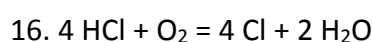
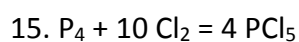
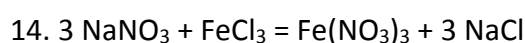
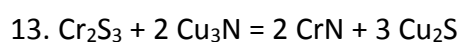
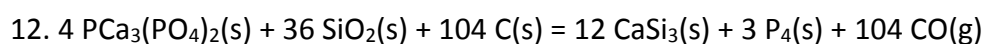
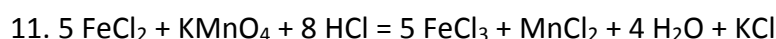
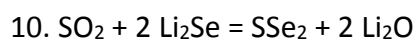
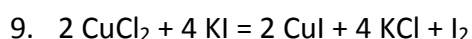
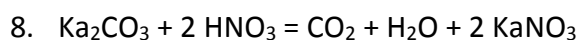
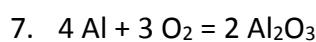
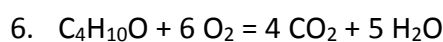
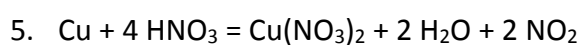
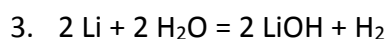
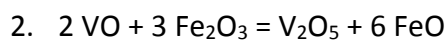
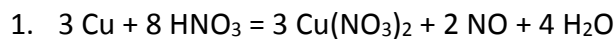
BALANCE THE GIVEN CHEMICAL EQUATIONS

Worksheet - 73

- $3 \text{ Cu} + \text{ ____ } \text{ HNO}_3 = 3 \text{ Cu(NO}_3)_2 + \text{ ____ } \text{ NO} + 4 \text{ H}_2\text{O}$
- $\text{ ____ } \text{ VO} + 3 \text{ Fe}_2\text{O}_3 = \text{ V}_2\text{O}_5 + \text{ ____ } \text{ FeO}$
- $2 \text{ Li} + \text{ ____ } \text{ H}_2\text{O} = \text{ ____ } \text{ LiOH} + \text{ H}_2$
- $\text{ ____ } \text{ K} + 2 \text{ H}_2\text{O} = \text{ ____ } \text{ KOH} + \text{ H}_2$
- $\text{ Cu} + \text{ ____ } \text{ HNO}_3 = \text{ Cu(NO}_3)_2 + \text{ ____ } \text{ H}_2\text{O} + 2 \text{ NO}_2$
- $\text{ C}_4\text{H}_{10}\text{O} + \text{ ____ } \text{ O}_2 = 4 \text{ CO}_2 + \text{ ____ } \text{ H}_2\text{O}$
- $\text{ ____ } \text{ Al} + 3 \text{ O}_2 = \text{ ____ } \text{ Al}_2\text{O}_3$
- $\text{ K}_2\text{CO}_3 + \text{ ____ } \text{ HNO}_3 = \text{ CO}_2 + \text{ H}_2\text{O} + \text{ ____ } \text{ KNO}_3$
- $\text{ ____ } \text{ CuCl}_2 + 4 \text{ KI} = \text{ ____ } \text{ CuI} + 4 \text{ KCl} + \text{ I}_2$
- $\text{ SO}_2 + \text{ ____ } \text{ Li}_2\text{Se} = \text{ SSe}_2 + \text{ ____ } \text{ Li}_2\text{O}$
- $\text{ ____ } \text{ FeCl}_2 + \text{ KMnO}_4 + 8 \text{ HCl} = \text{ ____ } \text{ FeCl}_3 + \text{ MnCl}_2 + 4 \text{ H}_2\text{O} + \text{ KCl}$
- $\text{ ____ } \text{ PCa}_3(\text{PO}_4)_2(\text{s}) + 36 \text{ SiO}_2(\text{s}) + 104 \text{ C}(\text{s}) = \text{ ____ } \text{ CaSi}_3(\text{s}) + 3 \text{ P}_4(\text{s}) + 104 \text{ CO}(\text{g})$
- $\text{ Cr}_2\text{S}_3 + \text{ ____ } \text{ Cu}_3\text{N} = 2 \text{ CrN} + \text{ ____ } \text{ Cu}_2\text{S}$
- $\text{ ____ } \text{ NaNO}_3 + \text{ FeCl}_3 = \text{ Fe(NO}_3)_3 + \text{ ____ } \text{ NaCl}$
- $\text{ P}_4 + \text{ ____ } \text{ Cl}_2 = \text{ ____ } \text{ PCl}_5$
- $\text{ ____ } \text{ HCl} + \text{ O}_2 = \text{ ____ } \text{ Cl} + 2 \text{ H}_2\text{O}$
- $3 \text{ Cl}_2 + \text{ ____ } \text{ NaOH} = 5 \text{ NaCl} + \text{ NaClO}_3 + \text{ ____ } \text{ H}_2\text{O}$
- $\text{ ____ } \text{ Li}(\text{s}) + \text{ N}_2(\text{g}) = \text{ ____ } \text{ Li}_3\text{N}(\text{s})$
- $3 \text{ F}_2 + \text{ ____ } \text{ Au} = \text{ ____ } \text{ F}_3\text{Au}$
- $\text{ ____ } \text{ H}_3\text{AsO}_4 + 5 \text{ H}_2\text{S} = \text{ As}_2\text{S}_3 + \text{ ____ } \text{ S} + 8 \text{ H}_2\text{O}$



ANSWERS





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