



## BALANCE THE GIVEN CHEMICAL EQUATIONS

### Worksheet - 76

1. \_\_\_\_  $\text{C}_8\text{H}_{18} + 25 \text{O}_2 = 16 \text{CO}_2 +$  \_\_\_\_  $\text{H}_2\text{O}$
2.  $3 \text{Na}_4(\text{SiO}_4)(\text{aq}) +$  \_\_\_\_  $\text{FeCl}_3(\text{aq}) = \text{Fe}_4(\text{SiO}_4)_3(\text{s}) +$  \_\_\_\_  $\text{NaCl}(\text{aq})$
3. \_\_\_\_  $\text{Cl} + 2 \text{KI} = \text{I}_2 +$  \_\_\_\_  $\text{KCl}$
4.  $3 \text{Ba}(\text{ClO}_3)_2 +$  \_\_\_\_  $\text{Na}_3(\text{PO}_4) = \text{Ba}_3(\text{PO}_4)_2 +$  \_\_\_\_  $\text{Na}(\text{ClO}_3)$
5.  $20 \text{KNO}_2 +$  \_\_\_\_  $\text{H}_2\text{O} = 10 \text{HNO}_3 +$  \_\_\_\_  $\text{KOHKNO}_2 + \text{H}_2\text{O}$
6.  $\text{Sr}(\text{OH})_2 +$  \_\_\_\_  $\text{HI} = \text{SrI}_2 +$  \_\_\_\_  $\text{H}_2\text{O}$
7.  $2 \text{Mg} +$  \_\_\_\_  $\text{HCl} =$  \_\_\_\_  $\text{MgCl} + \text{H}_2$
8.  $\text{Ba}(\text{OH})_2 +$  \_\_\_\_  $\text{HCl} = \text{BaCl}_2 +$  \_\_\_\_  $\text{H}_2\text{O}$
9.  $\text{P}_4 +$  \_\_\_\_  $\text{O}_2 =$  \_\_\_\_  $\text{P}_2\text{O}_3$
10. \_\_\_\_  $\text{As}_2\text{S}_3 + 28 \text{HNO}_3 + 4 \text{H}_2\text{O} =$  \_\_\_\_  $\text{H}_2\text{SO}_4 + 6 \text{H}_3\text{AsO}_4 + 28 \text{NO}$
11. \_\_\_\_  $\text{NaOH} + \text{AlBr}_3 =$  \_\_\_\_  $\text{NaBr} + \text{Al}(\text{OH})_3$
12. \_\_\_\_  $\text{HCl}(\text{aq}) + \text{O}_2(\text{g}) = 2 \text{H}_2\text{O}(\text{l}) +$  \_\_\_\_  $\text{Cl}_2(\text{g})$
13.  $2 \text{KMNO}_4 +$  \_\_\_\_  $\text{HCl} = 2 \text{KCl} + 2 \text{MnCl}_2 +$  \_\_\_\_  $\text{H}_2\text{O} + 5 \text{Cl}_2$
14. \_\_\_\_  $\text{LiOH} + \text{H}_2\text{SO}_4 = \text{Li}_2\text{SO}_4 +$  \_\_\_\_  $\text{H}_2\text{O}$
15.  $2 \text{Cr} +$  \_\_\_\_  $\text{Pb}(\text{NO}_3)_2 = 2 \text{Cr}(\text{NO}_3)_3 +$  \_\_\_\_  $\text{Pb}$
16.  $2 \text{Al} +$  \_\_\_\_  $\text{HBr} = 2 \text{AlBr}_3 +$  \_\_\_\_  $\text{H}_2$
17.  $\text{C}_{12}\text{H}_{22}\text{O}_{11} +$  \_\_\_\_  $\text{H}_2\text{SO}_4 = 12 \text{CO}_2 +$  \_\_\_\_  $\text{SO}_2 + 35 \text{H}_2\text{O}$
18. \_\_\_\_  $\text{KOH} + \text{H}_2\text{SO}_4 = \text{K}_2\text{SO}_4 +$  \_\_\_\_  $\text{H}_2\text{O}$
19.  $\text{Fe}_2\text{O}_3 + 3 \text{Cl}_2 +$  \_\_\_\_  $\text{KOH} = 2 \text{K}_2\text{FeO}_4 + 6 \text{KCl} +$  \_\_\_\_  $\text{H}_2\text{O}$
20. \_\_\_\_  $\text{CaH}_2 + 2 \text{H}_2\text{O} =$  \_\_\_\_  $\text{CaOH} + 3 \text{H}_2$



# ANSWERS

1.  $2 \text{C}_8\text{H}_{18} + 25 \text{O}_2 = 16 \text{CO}_2 + 18 \text{H}_2\text{O}$
2.  $3 \text{Na}_4(\text{SiO}_4)(\text{aq}) + 4 \text{FeCl}_3(\text{aq}) = \text{Fe}_4(\text{SiO}_4)_3(\text{s}) + 12 \text{NaCl}(\text{aq})$
3.  $2 \text{Cl} + 2 \text{KI} = \text{I}_2 + 2 \text{KCl}$
4.  $3 \text{Ba}(\text{ClO}_3)_2 + 2 \text{Na}_3(\text{PO}_4) = \text{Ba}_3(\text{PO}_4)_2 + 6 \text{Na}(\text{ClO}_3)$
5.  $20 \text{KNO}_2 + 20 \text{H}_2\text{O} = 10 \text{HNO}_3 + 10 \text{KOHKNO}_2 + \text{H}_2\text{O}$
6.  $\text{Sr}(\text{OH})_2 + 2 \text{HI} = \text{SrI}_2 + 2 \text{H}_2\text{O}$
7.  $2 \text{Mg} + 2 \text{HCl} = 2 \text{MgCl} + \text{H}_2$
8.  $\text{Ba}(\text{OH})_2 + 2 \text{HCl} = \text{BaCl}_2 + 2 \text{H}_2\text{O}$
9.  $\text{P}_4 + 3 \text{O}_2 = 2 \text{P}_2\text{O}_3$
10.  $3 \text{As}_2\text{S}_3 + 28 \text{HNO}_3 + 4 \text{H}_2\text{O} = 9 \text{H}_2\text{SO}_4 + 6 \text{H}_3\text{AsO}_4 + 28 \text{NO}$
11.  $3 \text{NaOH} + \text{AlBr}_3 = 3 \text{NaBr} + \text{Al}(\text{OH})_3$
12.  $4 \text{HCl}(\text{aq}) + \text{O}_2(\text{g}) = 2 \text{H}_2\text{O}(\text{l}) + 2 \text{Cl}_2(\text{g})$
13.  $2 \text{KMNO}_4 + 16 \text{HCl} = 2 \text{KCl} + 2 \text{MnCl}_2 + 8 \text{H}_2\text{O} + 5 \text{Cl}_2$
14.  $2 \text{LiOH} + \text{H}_2\text{SO}_4 = \text{Li}_2\text{SO}_4 + 2 \text{H}_2\text{O}$
15.  $2 \text{Cr} + 3 \text{Pb}(\text{NO}_3)_2 = 2 \text{Cr}(\text{NO}_3)_3 + 3 \text{Pb}$
16.  $2 \text{Al} + 6 \text{HBr} = 2 \text{AlBr}_3 + 3 \text{H}_2$
17.  $\text{C}_{12}\text{H}_{22}\text{O}_{11} + 24 \text{H}_2\text{SO}_4 = 12 \text{CO}_2 + 24 \text{SO}_2 + 35 \text{H}_2\text{O}$
18.  $2 \text{KOH} + \text{H}_2\text{SO}_4 = \text{K}_2\text{SO}_4 + 2 \text{H}_2\text{O}$
19.  $\text{Fe}_2\text{O}_3 + 3 \text{Cl}_2 + 10 \text{KOH} = 2 \text{K}_2\text{FeO}_4 + 6 \text{KCl} + 5 \text{H}_2\text{O}$
20.  $2 \text{CaH}_2 + 2 \text{H}_2\text{O} = 2 \text{CaOH} + 3 \text{H}_2$



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### **Further Questions?**

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