



BALANCE THE GIVEN CHEMICAL EQUATIONS

Worksheet - 78

- $2 \text{C}_2\text{H}_5\text{SH} + \text{O}_2 = 4 \text{CO}_2 + 6 \text{H}_2\text{O} + \text{SO}_2$
- $\text{KNO}_3(\text{aq}) + \text{H}_2\text{CO}(\text{aq}) = \text{K}_2\text{CO}_3(\text{s}) + \text{HNO}_2(\text{g})$
- $2 \text{Al}(\text{OH})_3 + \text{H}_2\text{SO}_4 = \text{Al}_2(\text{SO}_4)_3 + \text{H}_2\text{O}$
- $\text{KOH} + \text{H}_3\text{PO}_4 = \text{K}_3\text{PO}_4 + \text{H}_2\text{O}$
- $2 \text{NaCrO}_2 + 3 \text{Br}_2 + \text{NaOH} = 2 \text{Na}_2\text{CrO}_4 + \text{NaBr} + 4 \text{H}_2\text{O}$
- $2 \text{K}_3\text{AsO}_4 + \text{H}_2\text{S} = \text{As}_2\text{S}_5 + 6 \text{KOH} + \text{H}_2\text{O}$
- $\text{CuO} + 2 \text{NH}_3 = \text{N}_2 + \text{H}_2\text{O} + 3 \text{Cu}$
- $\text{Ba} + \text{CO}_2 = \text{BaO} + \text{C}$
- $\text{KClO}_3 = \text{KCl} + 3 \text{O}_2$
- $\text{PB}(\text{NO}_3)_2 = 2 \text{PBO} + \text{NO}_2 + \text{O}_2$
- $\text{Hg}(\text{BrO}_3)_2 + \text{NaC}_2\text{H}_3\text{O}_2 = \text{Hg}(\text{C}_2\text{H}_3\text{O}_2)_2 + \text{NaBrO}_3$
- $2 \text{C}_6\text{H}_{14} + \text{O}_2 = 12 \text{CO}_2 + \text{H}_2\text{O}$
- $\text{FeS}_2 + 12 \text{HCl} + 4 \text{NaCl} + 2 \text{H}_2\text{O} + 15 \text{O}_2 = \text{NaFeCl}_4 + 8 \text{H}_2\text{SO}_4$
- $\text{C}_2\text{O}_4^{2-} + \text{Ce}(\text{SO}_4)_2 = \text{CO}_2 + \text{Ce}_2(\text{SO}_4)_3 + \text{SO}_4^{2-}$
- $\text{Al} + \text{Fe}_3\text{N}_2 = \text{AlN} + 3 \text{Fe}$
- $4 \text{NH}_3 + \text{O}_2 = 4 \text{NO} + \text{H}_2\text{O}$
- $\text{NaOH}(\text{aq}) + \text{ZnCrO}_4(\text{aq}) = \text{Na}_2\text{CrO}_4(\text{aq}) + \text{Zn}(\text{OH})_2(\text{s})$
- $\text{Al}(\text{OH})_3 + \text{H}_2\text{SO}_4 = 2 \text{Al}(\text{OH})_3 + \text{H}_2\text{O} + \text{SO}_4$
- $2 \text{CrBr}_3 + \text{Na}_2\text{SiO}_3 = \text{Cr}_2(\text{SiO}_3)_3 + \text{NaBr}$
- $4 \text{C}_4\text{H}_9\text{O} + \text{O}_2 = 18 \text{H}_2\text{O} + \text{CO}_2$



ANSWERS

1. $2 \text{C}_2\text{H}_5\text{SH} + 9 \text{O}_2 = 4 \text{CO}_2 + 6 \text{H}_2\text{O} + 2 \text{SO}_2$
2. $2 \text{KNO}_3(\text{aq}) + \text{H}_2\text{CO}(\text{aq}) = \text{K}_2\text{CO}_3(\text{s}) + 2 \text{HNO}_2(\text{g})$
3. $2 \text{Al}(\text{OH})_3 + 3 \text{H}_2\text{SO}_4 = \text{Al}_2(\text{SO}_4)_3 + 6 \text{H}_2\text{O}$
4. $3 \text{KOH} + \text{H}_3\text{PO}_4 = \text{K}_3\text{PO}_4 + 3 \text{H}_2\text{O}$
5. $2 \text{NaCrO}_2 + 3 \text{Br}_2 + 8 \text{NaOH} = 2 \text{Na}_2\text{CrO}_4 + 6 \text{NaBr} + 4 \text{H}_2\text{O}$
6. $2 \text{K}_3\text{AsO}_4 + 5 \text{H}_2\text{S} = \text{As}_2\text{S}_5 + 6 \text{KOH} + 2 \text{H}_2\text{O}$
7. $3 \text{CuO} + 2 \text{NH}_3 = \text{N}_2 + 3 \text{H}_2\text{O} + 3 \text{Cu}$
8. $2 \text{Ba} + \text{CO}_2 = 2 \text{BaO} + \text{C}$
9. $2 \text{KClO}_3 = 2 \text{KCl} + 3 \text{O}_2$
10. $2 \text{PB}(\text{NO}_3)_2 = 2 \text{PBO} + 4 \text{NO}_2 + \text{O}_2$
11. $\text{Hg}(\text{BrO}_3)_2 + 2 \text{NaC}_2\text{H}_3\text{O}_2 = \text{Hg}(\text{C}_2\text{H}_3\text{O}_2)_2 + 2 \text{NaBrO}_3$
12. $2 \text{C}_6\text{H}_{14} + 19 \text{O}_2 = 12 \text{CO}_2 + 14 \text{H}_2\text{O}$
13. $4 \text{FeS}_2 + 12 \text{HCl} + 4 \text{NaCl} + 2 \text{H}_2\text{O} + 15 \text{O}_2 = 4 \text{NaFeCl}_4 + 8 \text{H}_2\text{SO}_4$
14. $\text{C}_2\text{O}_4^{\{-2\}} + 2 \text{Ce}(\text{SO}_4)_2 = 2 \text{CO}_2 + \text{Ce}_2(\text{SO}_4)_3 + \text{SO}_4^{\{-2\}}$
15. $2 \text{Al} + \text{Fe}_3\text{N}_2 = 2 \text{AlN} + 3 \text{Fe}$
16. $4 \text{NH}_3 + 5 \text{O}_2 = 4 \text{NO} + 6 \text{H}_2\text{O}$
17. $2 \text{NaOH}(\text{aq}) + \text{ZnCrO}_4(\text{aq}) = \text{Na}_2\text{CrO}_4(\text{aq}) + \text{Zn}(\text{OH})_2(\text{s})$
18. $2 \text{Al}(\text{OH})_4 + \text{H}_2\text{SO}_4 = 2 \text{Al}(\text{OH})_3 + 2 \text{H}_2\text{O} + \text{SO}_4$
19. $2 \text{CrBr}_3 + 3 \text{Na}_2\text{SiO}_3 = \text{Cr}_2(\text{SiO}_3)_3 + 6 \text{NaBr}$
20. $4 \text{C}_4\text{H}_9\text{O} + 23 \text{O}_2 = 18 \text{H}_2\text{O} + 16 \text{CO}_2$



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